

10-7-2011

Synthesis, Structural, Optical and Mechanical Characterization of SrB₂O₄ Nanorods

Rui Li

Lihong Bao

Xiaodong Li

University of South Carolina - Columbia, lixiao@cec.sc.edu

Follow this and additional works at: https://scholarcommons.sc.edu/emec_facpub



Part of the [Applied Mechanics Commons](#), and the [Other Mechanical Engineering Commons](#)

Publication Info

Published in *CrystEngComm*, Volume 13, Issue 19, 2011, pages 5858-5862.

©CrystEngComm 2011, Royal Society of Chemistry.

This article cannot be redistributed or further made available.

This article was first published by the Royal Society of Chemistry and can be found at <http://dx.doi.org/10.1039/C1CE05537B>

Li, R., Bao, L., & Li, X. (7 October 2011). Synthesis, Structural, Optical and Mechanical Characterization of SrB₂O₄ Nanorods. *CrystEngComm*, 13 (19), 5858 – 5862. <http://dx.doi.org/10.1039/C1CE05537B>

This Article is brought to you by the Mechanical Engineering, Department of at Scholar Commons. It has been accepted for inclusion in Faculty Publications by an authorized administrator of Scholar Commons. For more information, please contact digres@mailbox.sc.edu.

Cite this: *CrystEngComm*, 2011, **13**, 5858

www.rsc.org/crystengcomm

PAPER

Synthesis, structural, optical and mechanical characterization of SrB₂O₄ nanorods

Rui Li, Lihong Bao and Xiaodong Li*

Received 9th May 2011, Accepted 27th June 2011

DOI: 10.1039/c1ce05537b

Single crystalline strontium borate (SrB₂O₄) nanorods were synthesized for the first time *via* a facile sol-gel route at low temperature. The SrB₂O₄ nanorods have a good crystalline nature with the growth direction along the [511] orientation and they are transparent from the ultraviolet to the visible regimes. Nanoscale three-point bending tests were performed directly on individual nanorods to probe their mechanical properties using an atomic force microscope. The elastic modulus of SrB₂O₄ nanorods was measured to be 158.2 ± 2.8 GPa, exhibiting a significant increase compared with other borate nanostructures and bulk borates.

Introduction

One dimensional (1D) metal borate nanostructures have attracted tremendous attention due to their outstanding chemical inertness, high temperature stability, excellent mechanical properties, and low thermal expansion coefficient.^{1–8} As an important type of optical materials,⁹ strontium borates have shown various promising applications in high pressure sensors,¹⁰ magneto-optical media,¹¹ ultraviolet (UV) phosphors,¹² and detectors for thermoluminescent dosimetry.¹³ Guided by the SrO–B₂O₃ phase diagram,¹⁴ five strontium borate bulk materials have been synthesized, which are Sr₃B₂O₆, Sr₂B₂O₅, SrB₂O₄, SrB₄O₇, and Sr₂B₁₆O₂₆.¹⁵ The representative synthesis techniques include solid state reactions between strontium oxide and boric acid and the hydrothermal method. However, to our knowledge, strontium borate nanostructures are still not reported in the literature and their nanoscale-enabled properties are completely unknown.

In our previous work, 1D nanostructures of magnesium borate,¹ aluminium borate,² barium borate,³ and calcium borate⁴ have been successfully synthesized. These borate nanostructures exhibit superior mechanical properties. However, the mechanical properties of strontium borates are rarely reported, even for the bulk materials. On the other hand, the optical properties of strontium borate nanostructures have not been explored or theoretically predicted. In this paper, we report the facile synthesis, structural, optical and mechanical characterization of SrB₂O₄ nanorods. Single crystalline SrB₂O₄ nanorods were synthesized for the first time *via* a facile sol-gel route at a relatively low temperature of 800 °C. The X-ray diffraction (XRD) analysis revealed that the as-synthesized nanorods have an

orthorhombic phase of SrB₂O₄ with lattice constants of $a = 12.01$, $b = 4.34$ and $c = 6.59$ Å, and space group of *Pbcn*. The high-resolution transmission electron microscopy (HRTEM) characterization showed that individual SrB₂O₄ nanorods are single crystals with the growth direction along the [511] orientation. The ultraviolet-visible (UV-vis) absorption spectroscopy study unveiled that the SrB₂O₄ nanorods are transparent from the ultraviolet to the visible regimes. Atomic force microscopy (AFM) three-point bending tests were performed directly on individual nanorods to probe their mechanical properties. The elastic modulus of the SrB₂O₄ nanorods was measured to be 158.2 ± 2.8 GPa. Such structural, optical and mechanical information provides design guidelines for developing SrB₂O₄ nanostructure-based nanodevices and nanocomposites, and lays a constitutive foundation for modeling the nanostructures of strontium borates and other borate nanostructures.

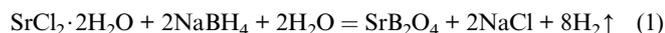
Experimental

All chemicals used in our experiments were purchased and used without any further purification. Strontium chloride dihydrate (SrCl₂·2H₂O) was purchased from Alfa Aesar. Sodium borohydride (NaBH₄) and cetyltrimethylammonium bromide (CTAB) were purchased from Sigma Aldrich.

In a typical experiment, 1 mmol of SrCl₂·2H₂O and 1.5 mmol of CTAB, and 6 mmol of NaBH₄ and 1 mmol of CTAB were separately dissolved in 12 ml distilled water with constant magnetic stirring at room temperature. Then the NaBH₄ solution was added dropwise to the SrCl₂ solution in 2 min. The reaction mixture was stirred for half an hour, and then stored in air at room temperature without stirring or shaking, and after 48 h a white gel was formed. The as-prepared precursors were placed in an alumina boat and annealed at 800 °C for 1 h in Ar. The white products were washed with distilled water and absolute ethanol three times, respectively, and dried at room temperature.

Department of Mechanical Engineering, University of South Carolina, 300 Main Street, Columbia, South Carolina, 29208, USA. E-mail: lixiao@cec.sc.edu; Fax: (+1) 803-777-0106; Tel: (+1) 803-777-8011

The chemical process for synthesizing SrB₂O₄ nanorods can be described in the following reaction:



The crystal structure of the products was characterized by XRD (Rigaku DMax 2200 using Cu K_α radiation). A few drops of ethanol solution containing the as-prepared products were deposited onto copper grids for transmission electron microscopy (TEM, Hitachi H-8000), and HRTEM (JEOL JEM 2010F) analyses. The solvent was then vaporized under ambient conditions. The resulting product was collected for characterization by scanning electron microscopy (SEM, Zeiss Ultra Plus FESEM). The UV-vis absorbance spectrum of the as-prepared products was obtained using a Beckman Coulter 640 DU spectrophotometer and quartz cuvettes from Starna. Chemical compositions and oxidation states of the products were determined by X-ray photoelectron spectroscopy (XPS, a Kratos Axis Ultra DLD instrument equipped with a monochromated Al K_α X-ray source and a hemispherical analyzer capable of an energy resolution of 0.5 eV). Thermogravimetric analysis (TGA) and differential thermal analysis (DTA) were carried out using a Thermal Analysis Instruments SDT2960 Thermogravimetric Analyzer with a scan rate of 20 °C per minute. Fourier transform infrared (FTIR) analysis was performed using a Perkin Elmer Spectrum 100 FTIR spectrometer fitted with a Diamond ATR attachment. Nanoscale three-point bending tests were performed directly on individual nanorods to probe their mechanical properties using AFM. To avoid sliding during the bending tests, both ends of the nanorods, which bridged over the trench, were clamped by electron beam induced deposition (EBID) of paraffin in SEM (FEI Quanta 200).^{16–18} This carbonaceous material thin layer and the strong adhesion force between the nanorod and the edges of the trench ensured the full clamping of the both ends of the suspended SrB₂O₄ nanorod during the bending tests, and thus the suspended SrB₂O₄ nanorod can be treated as a double clamped beam.

Results and discussion

Fig. 1a shows a representative XRD spectrum of the product synthesized *via* the sol-gel method at 800 °C for 1 h. All the peaks can be readily indexed to the orthorhombic phase of SrB₂O₄ (JCPDS file, No. 84-2175, with lattice constants of $a = 12.01$, $b = 4.34$ and $c = 6.59$ Å, and space group of *Pbcn*). No impurity peaks were observed, indicating high purity of the product. The SEM image of the product shown in Fig. 1b reveals that the product is composed of SrB₂O₄ nanorods with an average diameter of 80 nm and a length ranging from 1 to 3 μm. A high-magnification SEM image shows that the SrB₂O₄ nanorods have a circle-like cross-section (Fig. 1c). The energy dispersive spectroscopy (EDS, Fig. 1d) pattern of the SrB₂O₄ nanorods indicates that the nanorods contain Sr, B and O with the atomic ratio of 18.5 : 29.5 : 52.0, which is comparable to that obtained by XRD. The TEM image (Fig. 1e) shows that the SrB₂O₄ nanorods have very smooth surfaces and the average diameter and length of the nanorod are consistent with the SEM results. The representative HRTEM image (Fig. 1f) and the corresponding

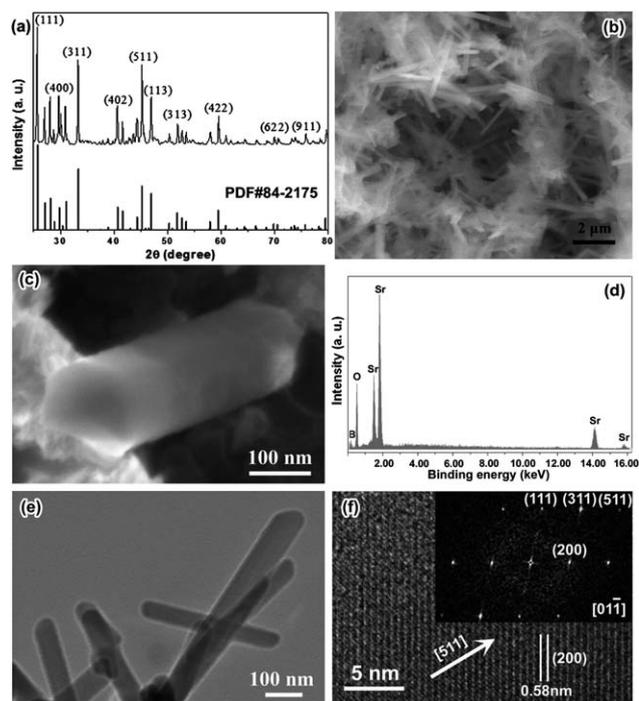


Fig. 1 (a) XRD pattern and the corresponding JCPDS standard pattern (PDF#84-2175), (b and c) SEM images, (d) EDS spectrum, (e) TEM image, (f) HRTEM image and (inset) corresponding FFT pattern of SrB₂O₄ nanorods.

fast-Fourier transform (FFT) pattern at the zone axis of [01 $\bar{1}$] jointly suggest that the synthesized SrB₂O₄ nanorods are single crystals with the growth direction along the [511] orientation, which is consistent with the stronger diffraction intensity of (511) in the XRD pattern in comparison with the standard XRD pattern in Fig. 1a. No defects such as dislocations and twins were found in individual SrB₂O₄ nanorods. The measured distance between the lattice fringes for the (200) planes is 0.58 nm, in good agreement with the XRD results.

To investigate the growth process of the SrB₂O₄ nanorods from the originating precursor (the white color gel obtained from the reaction mixture without heat treatment), both TGA and

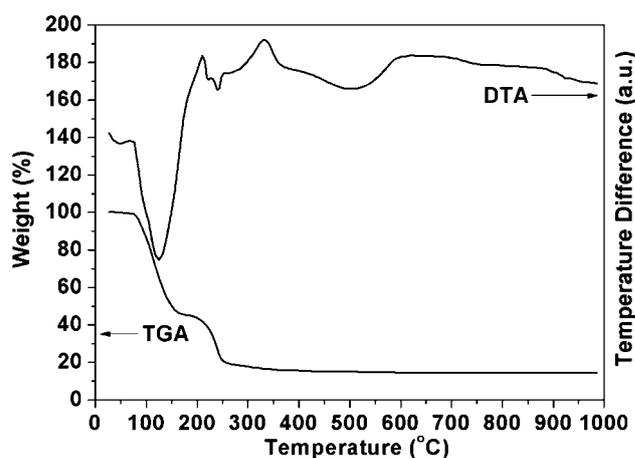


Fig. 2 TGA and DTA curves of the precursor.

DTA were performed, as is shown in Fig. 2. For the precursor, the weight loss was found to be about 82.3% in the range of 25–300 °C when heated in air, in good agreement with the weight loss from the removal of the adsorbed water molecules and the organic groups. Detailed analysis of the TGA curve (Fig. 2) reveals that the weight loss process occurred in two stages, the same as the thermal chemistry process of the reported hydrated borate.^{19,20} As determined by EDS, the white color gel (precursor) contains Sr, B, O, C, Na, Cl and Br elements. The contents of C and O reach up to 70 wt%, which indicates that the first stage is related to the removal of the adsorbed water molecules and the hydroxyl groups in the temperature range of 76–180 °C with a weight loss of 55.7%, and the second stage is ascribed to the decomposition of the organic components in the range of 200–300 °C with a weight loss of 26.1%. These two weight loss stages agree well with the two endothermic effects at about 125 °C and 240 °C in the DTA curve (Fig. 2). The exothermic effect at about 600–870 °C can be attributed to the crystallization of the SrB₂O₄ crystals.

XPS analysis was carried out to further characterize chemical compositions of the SrB₂O₄ nanorods (Fig. 3). The binding energies were corrected by taking C 1s at 284.8 eV. The survey spectrum, as shown in Fig. 3a, exhibits Sr 3d, B 1s, O 1s and C 1s

core levels, and O Auger peak. In Fig. 3b, the peaks at 133.2 eV and 134.8 eV in the curve can be attributed to Sr 3d_{5/2} and Sr 3d_{3/2}, respectively. In Fig. 3c, the peak at 191.8 eV can be assigned to B 1s core level. In Fig. 3d, the O 1s peak at 531.3 eV suggests that the oxygen exists as O²⁻ species in the SrB₂O₄ nanorods. The XPS results for the SrB₂O₄ nanorods are in good agreement with the reported values for bulk borates.^{3,21}

Fig. 4 shows the UV-vis absorption spectrum of SrB₂O₄ nanorods. It can be seen that the SrB₂O₄ nanorods are transparent from 200 nm (5.90 eV) to 1100 nm within the limitation of the instrument, whose working range is 200–1100 nm. The absorption spectrum is relatively featureless, and the broad and overall constant low absorbance from 200 to 1100 nm, including the nearly no absorbance from 200 to 300 nm, is consistent with the reported values for bulk SrB₂O₄ single crystals,²² demonstrating that the SrB₂O₄ nanorods are transparent from the ultraviolet to the visible regimes. This transparent character indicates that the SrB₂O₄ nanorods are very useful for the optical window application.

Fig. 5 shows the FTIR spectrum of the SrB₂O₄ nanorods. The spectrum exhibits broad absorptions in the 650–1600 cm⁻¹ range. Detailed FTIR analysis suggests that the SrB₂O₄ nanorods have all characteristics of bulk SrB₂O₄ crystals. The peak observed at

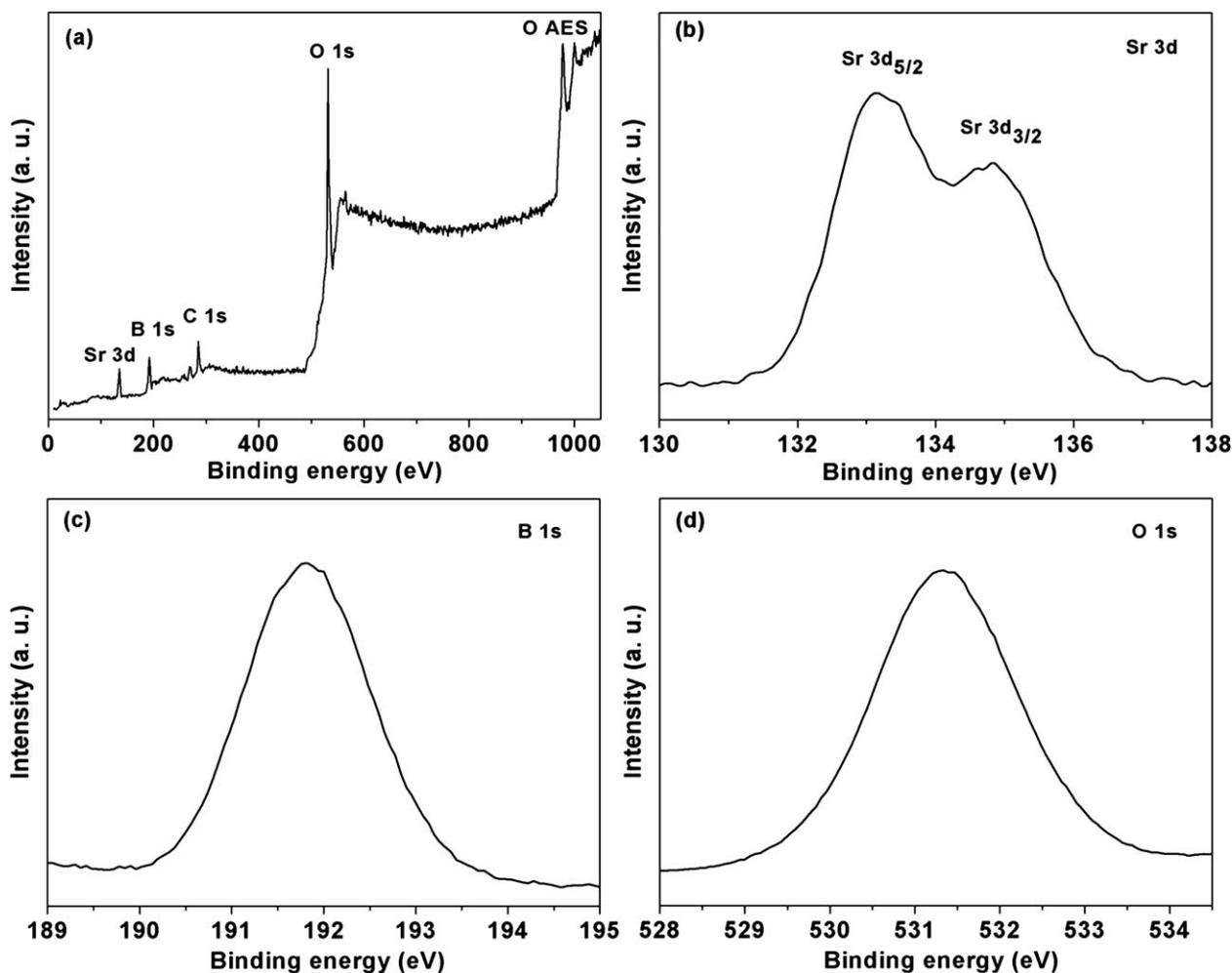


Fig. 3 XPS spectra of SrB₂O₄ nanorods. (a) Survey spectrum, (b)–(d) detailed spectra of Sr 3d, B 1s and O 1s core levels, respectively.

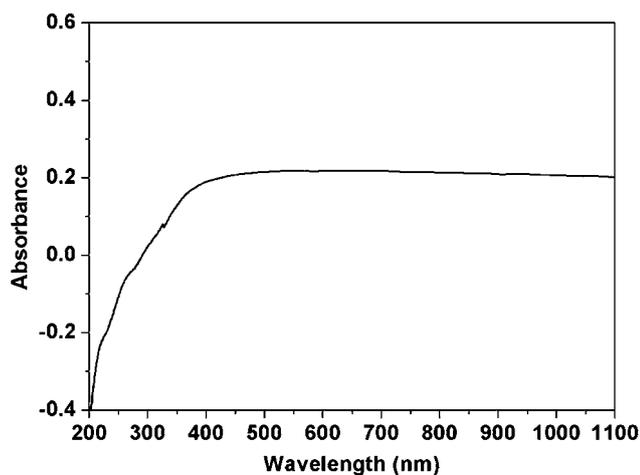


Fig. 4 UV-vis absorption spectrum of SrB₂O₄ nanorods.

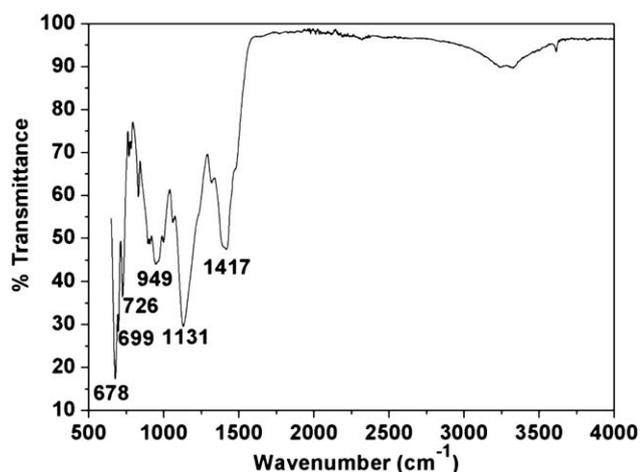


Fig. 5 FTIR spectrum of SrB₂O₄ nanorods.

949 cm⁻¹ is attributed to B–O extra-ring stretching.³ The bands at about 897 cm⁻¹ are due to the BO₃ symmetric stretching modes. The band at 1417 cm⁻¹ is due to the asymmetric stretching vibration of the group BO₃. The band at 1131 cm⁻¹ is assigned to the in-plane bending of the group BO₃. The bands observed at 726, 699, and 678 cm⁻¹ are due to the out-of-plane bending mode of the group BO₃. The FTIR spectrum confirms the existence of the trigonally coordinated boron atoms, consistent with the crystallographic study.²³

Nanoscale three-point bending tests^{17,24} were performed directly on individual SrB₂O₄ nanorods to probe their mechanical properties (Fig. 6a). To avoid sliding during the bending tests, both ends of the nanorod, which bridged over the trench, were clamped by EBID of paraffin.^{16–18} Fig. 6b and c are representative SEM and AFM images of an EBID-fixed nanorod, respectively. Fig. 6d shows representative force–piezo position (*F*–*Z*) curves obtained along the *Z*-direction on the nanorod directly sitting on a Si wafer and the suspended nanorod, respectively. The *F*–*Z* curves exhibit a strong linear relationship up to 450 nN. On the basis of elastic beam-bending theory, the elastic modulus of SrB₂O₄ nanorods, *E_n*, can be calculated from the equation:^{17,24}

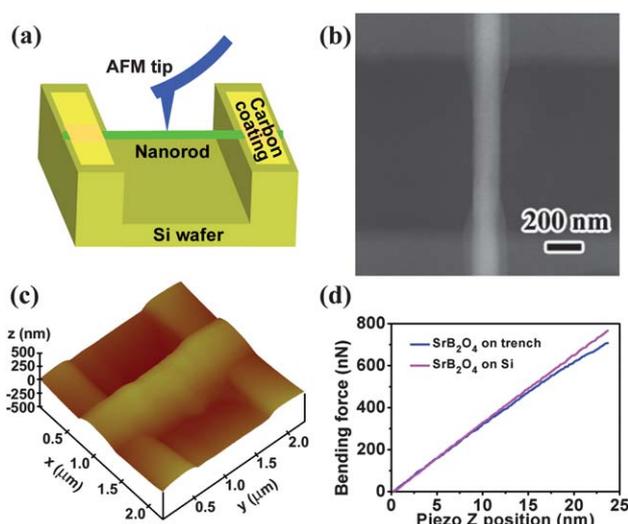


Fig. 6 AFM three-point bending test on an individual SrB₂O₄ nanorod. (a) Schematic image of an EBID-fixed SrB₂O₄ nanorod in a three-point bending test with an AFM tip. (b) SEM and (c) AFM images of a fixed SrB₂O₄ nanorod suspended over the trench. (d) Representative bending force–piezo position (*F*–*Z*) curves of the SrB₂O₄ nanorod directly sitting on Si wafer and the SrB₂O₄ nanorod bridging a trench with both ends of the nanorod fixed.

$$E_n = \frac{FL^3}{192d_n I} = \frac{k_n L^3}{192I} \quad (2)$$

where *I* is the moment of inertia. For a circle-shaped nanorod, *I* = π*r*⁴/4, where *r* is the radius of the nanorod. *L* is the suspended length of the nanorod and *F* is the applied load at its mid-point position. *k_n* (= *F*/*d_n*) is the spring constant of the nanorod, and can be obtained from the equation:²⁴

$$k_n = \frac{k_1 k_2}{k_1 - k_2} \quad (3)$$

where *k*₁ and *k*₂ are the slopes of the *F*–*Z* curves for the SrB₂O₄ nanorod directly sitting on the Si wafer and the suspended SrB₂O₄ nanorod, as shown in Fig. 6d, respectively. The average elastic modulus for the SrB₂O₄ nanorods was measured to be 158.2 ± 2.8 GPa. The mechanical properties of strontium borates have been rarely reported, even for the bulk materials. This is the first time to report the mechanical properties of strontium borates. Our SrB₂O₄ nanorods achieved 25.0%, 25.8%, 30.7%, 53.9%, and 15 720% increase in elastic modulus compared with Ca₂B₂O₅·H₂O nanobelts,⁴ Mg₂B₂O₅ nanowires,¹ Al₁₈B₄O₃₃ nanowires,² Al₄B₂O₉ nanowires,² and Ca₂B₂O₅ nanogrooves,⁴ respectively. The observed elastic modulus is also 1.8%, 60.8% increase compared with bulk borates: Mg₂B₂O₅,¹ Ca₂B₂O₅,²⁵ respectively; 60.5% decrease compared with bulk Al₁₈B₄O₃₃.²⁶ These findings are significant for designing strontium borate nanostructure-based nanocomposites and nanodevices, and lay a constitutive foundation for modeling the nanostructures of strontium borates and other borate nanostructures.

Conclusions

In summary, single crystalline strontium borate (SrB₂O₄) nanorods were synthesized for the first time *via* a facile sol–gel route at

low temperature. The SrB₂O₄ nanorods are transparent from the ultraviolet to the visible regimes. AFM-based nanoscale three-point bending tests were performed on individual SrB₂O₄ nanorods to probe their mechanical properties and the average bending elastic modulus is 158.2 ± 2.8 GPa, exhibiting a significant increase compared with other borate nanostructures and bulk borates.

Acknowledgements

Financial support for this study was provided by the National Science Foundation (CMMI-0653651 and CMMI-0824728), the ACS Petroleum Research Fund (ACS PRF 40450-AC10), and the University of South Carolina NanoCenter. We thank Douglas Blom and Soumitra Ghoshroy (the University of South Carolina EM Center) for TEM and HRTEM technical support.

References

- 1 X. Y. Tao and X. D. Li, *Nano Lett.*, 2008, **8**, 505–510.
- 2 X. Y. Tao, X. N. Wang and X. D. Li, *Nano Lett.*, 2007, **7**, 3172–3176.
- 3 R. Li, X. Y. Tao and X. D. Li, *J. Mater. Chem.*, 2009, **19**, 983–987.
- 4 L. H. Bao, Z. H. Xu, R. Li and X. D. Li, *Nano Lett.*, 2010, **10**, 255–262.
- 5 Y. Li, Z. Y. Fan, J. G. Lu and R. P. H. Chang, *Chem. Mater.*, 2004, **16**, 2512–2514.
- 6 R. Z. Ma, Y. Bando, T. Sato, C. C. Tang and F. F. Xu, *J. Am. Chem. Soc.*, 2002, **124**, 10668–10669.
- 7 C. C. Tang, E. M. Elssfah, J. Zhang and D. F. Chen, *Nanotechnology*, 2006, **17**, 2362–2367.
- 8 Y. C. Zhu, Y. Bando and R. Z. Ma, *Adv. Mater.*, 2003, **15**, 1377–1379.
- 9 T. I. Korshikova, S. V. Parkhomenko, A. V. Tolmachev, V. A. Tsurikov and R. P. Yavetskiy, *Inorg. Mater.*, 2008, **44**, 1345–1348.
- 10 F. Datchi, R. LeToullec and P. Loubeyre, *J. Appl. Phys.*, 1997, **81**, 3333–3339.
- 11 A. S. Aleksandrovsky, A. V. Malakhovskii, V. N. Zabluda, A. I. Zaitsev and A. V. Zamkov, *J. Phys. Chem. Solids*, 2006, **67**, 1908–1912.
- 12 D. S. Thakare, S. K. Omanwar, P. L. Muthal, S. M. Dhopte, V. K. Kondawar and S. V. Moharil, *Phys. Status Solidi A*, 2004, **201**, 574–581.
- 13 M. Santiago, A. Lavat, E. Caselli, M. Lester, L. J. Perisinotti, A. K. de Figuereido, F. Spano and F. Ortega, *Phys. Status Solidi A*, 1998, **167**, 233–236.
- 14 C. F. Chenot, *J. Am. Ceram. Soc.*, 1967, **50**, 117–118.
- 15 Z. H. Tang, X. A. Chen and M. Li, *Solid State Sci.*, 2008, **10**, 894–900.
- 16 W. Ding, D. A. Dikin, X. Chen, R. D. Piner, R. S. Ruoff, E. Zussman, X. Wang and X. Li, *J. Appl. Phys.*, 2005, **98**, 014905.
- 17 H. Ni, X. D. Li and H. S. Gao, *Appl. Phys. Lett.*, 2006, **88**, 043108.
- 18 H. Ni, X. D. Li, G. S. Cheng and R. Klie, *J. Mater. Res.*, 2006, **21**, 2882–2887.
- 19 Z. H. Liu and H. S. Huang, *Thermochim. Acta*, 2006, **448**, 59–62.
- 20 H. S. Huang and Z. H. Liu, *Thermochim. Acta*, 2007, **463**, 87–89.
- 21 G. D. Khattak, N. Tabet and M. A. Salim, *J. Electron Spectrosc. Relat. Phenom.*, 2003, **133**, 103–111.
- 22 J. B. Kim, K. S. Lee and J. N. Kim, *Mater. Lett.*, 1997, **31**, 93–97.
- 23 J. B. Kim, K. S. Lee, I. H. Suh, J. H. Lee, J. R. Park and Y. H. Shin, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 1996, **52**, 498–500.
- 24 X. Y. Tao, L. X. Dong, X. N. Wang, W. K. Zhang, B. J. Nelson and X. D. Li, *Adv. Mater.*, 2010, **22**, 2055–2059.
- 25 Manupriya, K. S. Thind, G. Sharma, V. Rajendran, K. Singh, A. V. G. Devi and S. Aravindan, *Phys. Status Solidi A*, 2006, **203**, 2356–2364.
- 26 S. C. Tjong and W. Jiang, *J. Appl. Polym. Sci.*, 1999, **73**, 2247–2253.